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Exothermic vs endothermic reactions worksheet

This is a FREE worksheet designed to accompany our Exothermic and Endothermal Reactions PowerPoint, which can be found for 4 pounds here. To watch the video, visit our youtube channel! Please leave us some feedback, we would appreciate it very much. 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Includes information on subject matter, chemistry, reactions and physical and chemical properties, physical change, precipitate, endothermal and exothermic reactions, chemical equations, reagents, products, mass preservation, open and closedPage 13Mic reactions NGSS activity package, 39 pages, information, activities, evaluations, questionnaire, vocabulary, teacher notes and guide. NGSS MATTER AND ITS MS-PS1 INTERACTIONS. chemistry, chemical reactions, atoms and elements, exothermic reactions, endothermal reactions, ws chemistry, chemical workpage 1471 Degrees mid-words and vocabulary illustrations including: Solid amorphous, Atom, Atomic Theory, Boiling Point, Linking, Chemical Change, Chemical Formula, Chemical Property, Classification, Composite, Condensation, Conduction, Convection, Crystal Solid, Density, Electron, Electronic Cloud, ElePage 156th, 7th, 8th, 9th, 10th, 11th, 12th, Higher Education, HomeschoolPage 16Is a Power Presentation Point in the matter. 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This transfer is in the form of thermal energy, i.e. heat. A chemical reaction can be exothermic or endothermic. Exothermic reactions Exo- means 'out' - to think of words like 'exit'. Exothermic reactions release heat into your environment. The image shows an example of an exothermic reaction: Clearly, this looks like an explosion. Explosions are exothermic reactions, but a reaction can be exothermic without exploding. Any temperature increase indicates an exothermic reaction. If you had a container when an exothermic reaction occurred, you would feel the heat, and a thermometer inside the vessel would show an increase in temperature. Examples of exothermic reactions are combustion (burning) and neutralization. Hand warm-up pads contain chemicals that undergo an exothermic reaction when pressed together. Endothermal reactions Endo- means 'inside' - an endoscope is a camera that a doctor puts inside a patient's body. Endothermal reactions take energy from your environment. If you had a container when an endothermic reaction occurred, you'd feel your hands cool down, as the reaction would require heat from the vessel, and even your hands, if you kept holding it! A thermometer in the container would show a decrease in temperature. The energy of the products is greater than the energy of the reagents. Examples of endothermal reactions include electrolysis, photosynthesis, and thermal decomposition. In all these reactions, we need to keep putting energy into the reaction to keep it going. In electrolysis the power source is the battery, in photosynthesis the energy is light, and in thermal decomposition the energy source is a source of heat (like a flame). Cold ear pads to relieve injuries can use chemicals that suffer an endothermic reaction. Reaction profiles A reaction profile is a kind of graph that compares the energy of a set of chemicals before, during and after a chemical reaction. Before the reaction is to the left, and after the reaction is on the right. The energy is on the vertical axis - the high energy is high and the low energy is low. In an exothermic reaction (on the left), the energy of the chemicals starts high and ends lower. The power difference is released into the environment. In an endothermic reaction (on the right), chemical energy starts low and ends higher. The power difference is taken from the environment. Both reaction profiles have a spike in the middle, where the energy of the chemicals is higher than all. This is because almost all chemical reactions some energy input so they can start to happen. The energy needed to initiate a chemical reaction is called activation energy - it is the energy needed to make the reaction active. Reaction profiles help us visualize the difference between exothermic and endothermal reactions, and compare reactions involving different amounts of energy - for example, an explosion releases a lot of energy in a short time. Now it's time for some questions about this. 5th, 6th, 7th, 8th, 9th, 10th, 11th, 12th, HomeschoolPage 26, 7th, 8th, 8th, 10th, 11th, 12th, HomeschoolPage 36th, 7th, 8th, 9th, 10th, 11th, 12th, HomePageschool 46th, 7th, 8th, 9th, 10th, 11th, 12th, HomeschoolPage 55, 6th, 7th, 8th, 9th, 10th, 11th, 12th, Homeschool, StaffPage 65th, 6th, 7th, 8th, 9th, 10th, 11th, 12th, Higher Education, Adult Education, Home School, PersonalPage 75, 6th, 7th , 8th , 9th, 10th, 11th, 12thPage 86, 7th, 8th, 9th, 10th, 11th, 12th, Higher Education, HomeschoolPage 93a, 4a, 5th, 6th, 7th, 8th, 9th, 10th, 11th, 12thPage 108, 9th, 10th, 11th, 12th, Higher Education, Adult Education, Home School, Personal, Non-Specific GradePage 119, 10th, 11th, 12th, Higher Education, Adult Education, Homeschool, StaffPage 12 This is a Power Point presentation on chemical reactions. 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